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~~$\ln K = (\Delta rS/R) - (\Delta rH/RT)$~~  The equilibrium for the dissociation of one mole of iodine vapour is  $3.94 \times 10^{-4}$  at 900K and the standard enthalpy is 154 kJmol<sup>-1</sup>. Calculate the equilibrium constant at 1200K assuming the standard enthalpy remains constant.  ~~$\ln (K_2/K_1) = - \Delta rH/R (1/T_2 - 1/T_1)$~~  , ANS =  $6.7 \times 10^{-2}$  at 1200K.

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### ~~Physical Chemistry—American Chemical Society~~

Preparing for Your ACS Examination in Physical Chemistry: The Official Guide by American Chemical Society published by Chem Ed Exams (2009)

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### ~~DEPARTMENT OF CHEMISTRY AND BIOCHEMISTRY~~

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