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Equations and Formulas of Chemical Reaction Engineering

Type or Name General Formula 1. Alkane or paraffin (also saturated hydrocarbons) 2. Alkene or olefin (unsaturated hydrocarbons) 3. Alkyne 4. Alcohol 5. Ether 6. Aldehyde 7. Ketone 8. Carboxylic Acid 9. Grignard reagent 10. Acyl halide 7 13282AICEtext 4/12/04 12:20 PM Page 7

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iii PREFACE The Fundamentals of Engineering (FE) Supplied-Reference Handbook is the only reference material allowed in the FE examination. Many examinees find that it is helpful to review this book before exam day to become familiar with the reference material it contains.

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Chemical Engineering Formula Sheet where reaction is in solution, ([] refers to molarity) where reaction is in the gas phase, (p partial pressure) Gram equivalent weight— A. (nonredox reaction) the mass in grams of a substance equivalent to 1 gram-atom of hydrogen, 0.5 gram- atom of oxygen, or 1 gram-ion of the hydroxyl ion.

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The Fundamentals of Engineering (FE) exam is generally your first step in the process to becoming a professional licensed engineer (P.E.). It is designed for recent graduates and students who are close to finishing an undergraduate engineering degree from an EAC/ABET-accredited program.

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NCEES FE exam information

This article will help you improve your Excel spreadsheet skills and will provide you with tools to solve common chemical engineering problems. Ever since Lotus 1-2-3 and the IBM PC became popular in the early 1980s, chemical engineers have been using spreadsheets for day-to-day problem-solving.

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This list gives you some of the most common conversion factors you need in thermodynamics. Acceleration: $1 \text{ m/s}^2 = 100 \text{ cm/s}^2$. Area: $1 \text{ m}^2 = 10^4 \text{ cm}^2 = 10^6 \text{ mm}^2$. Density: $1 \text{ g/cm}^3 = 1 \text{ kg/L} = 1,000 \text{ kg/m}^3$. Energy, heat, work, internal energy, enthalpy: $1 \text{ kJ} = 1,000 \text{ J} = 1,000 \text{ N} \cdot \text{m} = 1 \text{ kPa} \cdot \text{m}^3$. $1 \text{ kJ/kg} = 1,000 \text{ m}^2/\text{s}^2$.

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