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Page 29 and 30: 29

Information: Average Atomic

Mass. Page 31 and 32: 31

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Page 33 and 34: 33 An  
electron that absorbs energy  
. Page 35 and 36: 35  
Information: Orbitals So far  
we ; Page 37 and 38: 37  
ChemQuest 12 Name: \_\_\_\_\_  
Page 39 and 40: 39 Critical  
Thinking Questions 6. H

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Structures tr I Name Date:  
Hour: Introduction Questions  
Draw the Lewis structures  
for the following molecules.  
Part c is done for you. a)  
H<sub>2</sub>O (oxygen is central) b)  
SO<sub>2</sub> c) N<sub>2</sub> It t-t (a [Jrt @  
N=N @ @ f\ @ 2 In question 1  
above you should see three  
different kinds of bonds

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ofbonds as best you can.

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Structunes tr I Name Date:

Hour: Introduction Questions

Draw the Lewis structures

for the following molecules.

Part c is done for you. a)

H<sub>2</sub>O (oxygen is central) b)

SO<sub>2</sub> c) Nz It t-t (a [Jrt @

N=N @ @ f\ @ 2 In question 1

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different kinds ofbonds

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ChemQuest 37 SslgDcll?g

E"alue Date Hour I

ntroduction Questions

Information: Subscripts and

Coefficients A subscript is

a small number that tells

you how many atoms are in a

compound. For example, in

CaCl<sub>2</sub> the two is the

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## Answers

subscript and it tells us that there are two chloride ions bonded to one calcium. A coefficient tells also tells us how many atoms or compounds there are, but in a different way.

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four times your answer to  
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electrons pictured in the  
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