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Entropy, and Gibbs Free Energy AMIT KUMAR, ME, AIR 1,

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Cheat Sheet

This list gives you some of the most common conversion factors

you need in thermodynamics. Acceleration: $1 \text{ m/s}^2 = 100 \text{ cm/s}^2$.

Area: $1 \text{ m}^2 = 10^4 \text{ cm}^2 = 10^6 \text{ mm}^2$. Density: $1 \text{ g/cm}^3 = 1 \text{ kg/L} =$

$1,000 \text{ kg/m}^3$. Energy, heat, work, internal energy, enthalpy: $1 \text{ kJ} =$

$1,000 \text{ J} = 1,000 \text{ N}\cdot\text{m} = 1 \text{ kPa}\cdot\text{m}^3$. $1 \text{ kJ/kg} = 1,000 \text{ m}^2/\text{s}^2$.

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Basic Thermodynamic Formulas (Exam Equation Sheet) Control Mass (no mass flow across system boundaries) Conservation of mass: $\dot{m} = \dot{m}_{in} - \dot{m}_{out}$. Conservation of energy (1st Law): $\dot{Q} = \dot{m} \Delta h = \dot{m} \Delta u + \dot{W}$

~~Basic Thermodynamic Formulas (Exam Equation Sheet)~~

my thermodynamics cheat sheets - 12000.org Thermodynamics key facts (7/9) • Ideal gas law • 1. st. form : $pV = nRT$. p = Pressure, V = Volume, n = number of molecules, R = Boltzmann's constant, T = temperature [in K] • 2. nd. form : $p = \frac{1}{3} n m \bar{v}^2$

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Thermodynamics key facts (3/9) • Heat energy needed to change phase • The latent heat h_{fg} determines the energy needed to change the phase of a mass m . • Units of h_{fg} will be kJ/kg . 1. - can be fusion or vaporization • This energy is either absorbed (solid \rightarrow liquid \rightarrow gas) or released (gas \rightarrow liquid \rightarrow solid) • A phase change takes place at constant temperature.

Revision : Thermodynamics

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didn't let us use anything in our test other than a pencil. Calculator and the thermodynamic property table book. No cheat sheet, reference sheet. Nothing.

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Thermodynamics ? $T A v = A^2 v P =$ rate of heat transfer $Q =$ thermal energy $A =$ Area of thermal conductivity $U =$ coefficient of heat conductivity (U-factor) $T =$ change in temperature $R =$ resistance to heat flow (R-value) $v =$ velocity net = net power radiated = $5.6696 \times 10^{-8} 1$, $T 2 =$ temperature at time 1, time 2 $v =$ flow velocity

~~Engineering Formula Sheet~~

IV. The First Law of Thermodynamics as a Rate Equation A. Most general form $dE_{c.v.} / dt = \dot{Q} \text{ ? } c.v. + \dot{W} \text{ ? } c.v. + \dot{m} \text{ ? } e_{in} - \dot{m} \text{ ? } e_{out}$ out rate of change of energy in c.v. = rate of heat added to c.v. rate of work done + rate of energy flow in to c.v. rate of energy flow out of c.v. B. For a steady flow process $d/dt = 0$ and $\dot{m} \text{ ? } in = \dot{m} \text{ ? } out = \dot{m}$? $\dot{Q} \text{ ? } c.v. + \dot{W} \text{ ?}$

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Thermodynamics: the study of energy, energy transformations and its relation to matter. The analysis of thermal systems is achieved through the application of the governing conservation equations, namely Conservation of Mass, Conservation of Energy (1st law of thermodynamics), the 2nd law of thermodynamics and the property relations.

~~Basic Concepts of Thermodynamics~~

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engineering work, pressures are often measured with respect to atmospheric pressure rather than with respect to absolute vacuum. $P_{abs} = P_{atm} + P_{gauge}$ In SI units the derived unit for pressure is the Pascal (Pa), where $1 \text{ Pa} = 1 \text{ N/m}^2$. This is very small for engineering purposes, so usually pressures are given in terms of kiloPascals ($1 \text{ kPa} = 10^3 \text{ Pa}$),

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