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Answer: Calculate the number of half-lives; $0.003 \text{ seconds} \times 1 \text{ half-life} = 3 \text{ half-lives}$
 $0.001 \text{ second} \square$ After 0 half-lives, 10

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g are left. After 1 half-life, 5 g are left.
After 2 half-lives, 2.5 g are left. After 3
half-lives, 1.25 g are left.

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HALF-LIFE WORKSHEET 1. What is radioactivity? 2. What is half-life? ... Use the following graph to answer questions 7-10. 7. How long is a half-life for carbon-14? _____ 8. If only 25% of the

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carbon-14 remains, how old is the material containing the carbon-14?_____ 9. If a sample originally had 120 atoms of carbon-14, how many atoms will remain ...

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Reference Table on side to assist you in answering the following questions.

Equations: Yt lifes: As-81 = 33 seconds

Au-198 = 2.69 days

HALF-LIFE PROBLEMS

If each half life is 5 seconds, then in one

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minute. (60 seconds) there are 12 half lives. Therefore the answer is: 2)

Iodine-131 has a half life of 8 days. What fraction of the original sample would remain at the end of 32 days? Using the same fraction, you must figure out n . If the half life is 8 days,

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would remain at the end of 32 days? 32 days / 8 days = 4 half lives. $1 \left(\frac{1}{2} \left(\frac{1}{4} \left(\frac{1}{8} \left(\frac{1}{16} \right. \right. \right. \right. \right.$ will remain.

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Answers □ Kids ...

This worksheet allows students to practice solving all types of problems involving half-life. A complete answer key is provided at the end. While this worksheet was designed for an honors-level Chemistry class, it can easily be modified for lower ability students.

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Half Life Worksheet | Teachers Pay Teachers

Problem #3: Os-182 has a half-life of 21.5 hours. How many grams of a 10.0 gram sample would have decayed after exactly three half-lives? Solution: $(1/2)^3 = 0.125$

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(the amount remaining after 3 half-lives)

$10.0 \text{ g} \times 0.125 = 1.25 \text{ g}$ remain $10.0 \text{ g} -$

$1.25 \text{ g} = 8.75 \text{ g}$ have decayed Note that the length of the half-life played no role in this calculation.

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