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150. $\text{Ag}_2\text{CrO}_4(\text{s}) + 2 \text{KNO}_3(\text{aq})$ 2. How many mL of 0.

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silver nitrate are added to 100. mL of
0.400 M potassium chromate? 2

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0.150 L AgNO_3

0.500 moles AgNO_3 1 moles Ag_2CrO_4

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As we learned previously, double
replacement reactions involve the
reaction between ionic compounds in

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156 solution and, in the course of the reaction, the ions in the two reacting compounds are “switched” (they replace each other). Because these reactions occur in aqueous solution, we can use the concept of molarity to directly calculate the number of moles of reactants or products that will ...

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